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laboratory. 2. To learn the capabilities of the different types of balances which are available in the laboratory. 3. To relate the concept of significant numbers to the accuracy of mass and volume measurements. PRINCIPLES: One of the most important operations in a chemistry laboratory is the massing of

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Pre-lab Assignment/Questions *Note- this pre-lab must be finished before you come to lab. 1. Calculate the molar mass of citric acid C 6 H 8 O 7. 2. Calculate the volume of 2.5 M NaOH needed to make 200 mL of 0.04 M NaOH. 3. The balanced equation for KHP and NaOH is given in the introduction to this lab.

Lab 7 Worksheet | College Chemistry 1 Labs

CH4(g) +O2(g) → CO2(g) +2H3(g)O ΔH= −890.4 kJ CH 4 (g) + O 2 (g) → CO 2 (g) + 2H 3 (g) O Δ H = − 890.4 kJ. The negative sign serves to reinforce the fact that heat is produced rather than absorbed or that this is an exothermic reaction. We can use stoichiometry to calculate the enthalpy of any amount of a reactant.

Lab 9 Introduction | Chemistry I Laboratory Manual

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